

Metallic solutions and calorimetry

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Abstract. In materials science, accurate high temperature thermodynamic measurements are needed to predict and to understand the properties of n-components systems. Two examples have been chosen to highlight this problem : i) from molar heat capacity determinations, the miscibility gap of the solid HgTe-CdTe system has been found ; ii) in spite of several empirical and theoretical models, measurements of the enthalpy of formation of Cd-Ga-In-Sn-Zn liquid alloys are unavoidable. These examples show the necessity to develop very precise calorimetric methods in the high temperature range.

INTRODUCTION

For high temperature science, thermodynamics is a powerful tool which allows to understand the properties of materials (alloys, ceramics,...). From results obtained at low temperature we can predict what it will happen at high temperature : that is the main interest of this science. Moreover taking into account the difficulties encountered to perform high temperature measurements, data predicted by using the well known thermodynamic relations seem more reliable and more accurate than those we could hope to determine.

But fairly often, in materials science, high temperature determinations are unavoidable. Usually it is uneasy and often very expensive to work in such conditions. Thus the choice of the experimental technique and of the decisive measurement is the essential point.

In high temperature range, all techniques based on the study of thermal effects developed during experiments (calorimetry, thermal analysis) are efficient. So, from some examples, the main purpose of this lecture is to highlight the interest of accurate thermal measurements at high temperature in materials science.

METHODS

In the field of high temperature thermochemistry, for systems containing metals, oxides or salts, Gibbs free energy measurements are very scarce. Generally the most often measured functions are enthalpies of formation, enthalpies of transition, equilibrium temperatures, heat capacities,...So, the main techniques used are thermal analysis and calorimetry. The number of apparatus (calorimeters and thermal analysers) which have been designed to meet requirements for the best solution of an actual problem is considerable. But these two techniques and apparatus have been described in several text-books and reviews. The choice of an apparatus depends on a large number of factors : the experimental temperature range, the optimal accuracy, the size of the thermal effect, the reactivity of samples,...The difficulties increase considerably with temperature ; for these two techniques, the common difficulties arise from the choice of containers (to avoid contamination), the vaporization of materials, the calibration procedures, the heat leaks by conduction, radiation,...

The specific difficulties encountered up to 1000K in differential thermal analysis and the solutions adopted have been recently described by J.L. Jorda in : "Analyses thermiques à haute température : mode ou nécessité" (ref.1).

Several descriptions of calorimeters used in metallurgy have been published (ref. 2 and 3). It seems that the majority of the high temperature apparatus belongs to the classes of isoperibol high temperature calorimeters and of conduction high temperature calorimeters. Unfortunately, we have to regret a decline in adiabatic high temperature calorimetry which has been a considerable source of data in metallurgy. With isoperibol high temperature calorimeters, Ferro and coworkers (ref. 4), Predel and coworkers (ref. 3) carried out a large number of enthalpies of formation of solid and liquid alloys. The Calvet high temperature, operable on the temperature range 500-1300 K is used by several groups or scientists (R. Castanet, C. Colinet, M. Gambino, E. Hayer, O.J. Kleppa, Z. Moser,... publications of whom are well-known). Up to 1300 K, calorimeters available to determine enthalpies of formation of alloys are not numerous. A very-high temperature operable until 1800 K is used by M. Gaune-Escard, J. Hertz, J.L. Jorda, A. Navrotsky,...

Nowaday, a new class of calorimeter, the differential scanning calorimeters (D.S.C.) allows to obtain quickly the molar heat capacities and the enthalpies of transition. Unfortunately, the upper temperature of these apparatus is about 1000 K.

APPLICATIONS

To highlight the advantage of precise high temperature calorimetric measurements, two examples have been chosen ; they concern the determination of : i) the miscibility gap of CdTe-HgTe solid alloy, ii) the heat of formation of the Cd-Ga-In-Sn-Zn liquid alloy.

CdTe-HgTe system

Two experimental ways (calorimetric and potentiometric) were followed to investigate the CdTe-HgTe solid phase region.

Much has been published about CdTe-HgTe system ; we will not report here, in details, the thermodynamic data but we would recall that the equilibrium phase diagram is of the simplest type : HgTe and CdTe form a complete series of solid solutions which have the ZnS structure. The Hg and Cd atoms are randomly distributed over the sites of one sublattice. Among the many thermodynamic data published, the heat capacities of CdTe-HgTe alloys are missing. Unfortunately, for a better phase diagram assessment, the values of molar heat capacity of components and alloys versus temperature are required. Moreover data concerning the solid phase are very scarce.

Preparation and analysis of alloys. Well known mass of tellurium, cadmium (in powder) and mercury of high purity (99.999 at. %) were placed in quartz crucibles (10 mm in diameter) sealed under vacuum (10^{-3} Pa) and heated at 673 K. Every three hours, the temperature of the furnace was increased of 50 K. At 723 K, the temperature was stabilized during one week then slowly decreased in one day. Ingots so obtained were ground. Using a special mould, pellets were prepared (14 mm in diameter, 10 mm high). These pellets were maintained under vacuum, in sealed ampoule, at 723 K, during one week then slowly cooled until 523 K. All these alloys were analysed by X-ray diffraction using the $\text{CuK}\alpha$ radiation with a DRON apparatus. The variation of the lattice parameter ($a/\text{\AA}$) of the CdTe-HgTe solid solutions follows the Vegard's law and is given by :

$$a/\text{\AA} = 6,4637 + 2,02 \cdot 10^{-2} x_{\text{CdTe}} \quad \text{with a standard deviation } 2S_0 = \pm 0,002 \text{\AA}.$$

1. C_p measurements

Ideally, determination of molar heat capacities should be made with an adiabatic calorimeter. But, along with the recent progress achieved in differential scanning calorimetry, it seems quite natural to employ this technique to carry out C_p values of multicomponent systems on large temperature ranges and at several compositions in order to detect the critical lines. So, using a differential scanning calorimeter, C_p values versus temperature of CdTe-HgTe solid alloys have been carried out.

The specific heat of a sample is calculated from the relationship :

$$C_p = (dH/dT)_p$$

and if $T=f(t)$

$$((dH/dt)/(dT/dt))_p \text{ then } (dH/dt)_p = C_p(dT/dt)$$

where H, T and t are respectively enthalpy, temperature and time.

If the temperature increases from T_1 to T_2 during the period t_2-t_1 , the equation may be written :

$$\int (dH/dt)_p = \int C_p(dT/dt)$$

with the assumption that the heat capacity is constant on a narrow temperature range,

then

$$C_p = [H]_{t_1}^{t_2} / [T]_{t_1}^{t_2}$$

Consequently, the practical determination of $C_p = f(T)$ with a differential scanning calorimeter demands the knowledge of the temperature increase of the calorimeter containing the experimental and the reference cells and the variation of the enthalpy during the same period.

The apparatus (DSC 111- Setaram company) is designed as a Calvet calorimeter with two cylindrical cells (reference and laboratory cells). It should be emphasized that two calibration methods have been employed : in one method, a standard NIST alumina crystal (ref. 5) is used for comparison and, in the other, direct Joule-heating of an electrical resistance located in the cell is used. Such an apparatus can be operated between 173 K and 1023 K.

Consequently, in measuring heat capacities, the following procedure was adopted ; with the same experimental conditions (constant heat rate, identical scanned temperature range, identical argon flow,...) :

- i) the "zero test", with two identical empty containers, allows to determine the thermal disequilibrium between the reference and laboratory cells,
- ii) the " experimental test" performs with the sample in the experimental cell,
- iii) the " calibration test" obtained by using either the Joule effect or a well known mass of standard NIST alumina crystal. The discrepancy between our experimental results and the reference data (NIST) was always less than 0.5%.

Great care has to be taken in protecting samples in the cell : alloys (about 200 mg) were placed in gastight stainless crucibles (6 mm diameter , 20 mm length) and all experiments were performed under purified argon flow. On the temperature range 300-523 K, every five Kelvins, the heat capacity of the sample have been determined. The heating rate is 1.5 K.mn⁻¹ during 200 seconds and the temperature is maintained constant during 400 seconds. Temperatures were determined at $\pm 0.5^\circ\text{C}$ and molar heat capacities were measured with an accuracy of about 2%.

Over the temperature range 300 K-523 K, the molar heat capacity of CdTe, HgTe and of ten HgTe-CdTe alloys (with $x_{\text{CdTe}} = 0.0996, 0.118, 0.253, 0.350, 0.400, 0.500, 0.544, 0.600, 0.700$ and 0.800) has been determined.

Molar heat capacity of CdTe and HgTe. The molar heat capacity of these two compounds (CdTe and HgTe) may be expressed as

$$C_p, \text{CdTe} = 23.902 + 0.00576 T \quad \text{in J.K}^{-1}.\text{mol}^{-1} \quad \{1\}$$

$$C_p, \text{HgTe} = 25.445 + 0.004068 T \quad \text{in J.K}^{-1}.\text{mol}^{-1} \quad \{2\}$$

A good agreement is remarked with values obtained by using Kopp-Neuman rule ; the heat capacities of elements (Te, Cd) are from Hultgren et al. compilation (ref. 6). These results have been compared to those carried out by Malkova (ref. 7) and Kelemen (ref. 8) for CdTe and HgTe respectively.

Molar heat capacity of CdTe-HgTe alloys. The experimental values of C_p of the CdTe-HgTe system have been integrally published (ref. 9). Using the equations {1} and {2} and Neumann-Kopp rule, the molar heat capacities have been also calculated.

From these results it may be noticed that :

i) if $x_{\text{CdTe}} < 0.253$, the $C_p = f(T)$ law is quasi-linear and reasonably reliable estimates are obtained with Neumann and Kopp rule (Fig. 1)

ii) if $x_{\text{CdTe}} > 0.253$, $C_p = f(T)$ plots (Fig. 2) exhibit a step at the temperature T_s with a linear variation of C_p if $T > T_s$. Below the temperature of the step, the variation of C_p is irregular and $\Delta C_p = (C_{p,\text{exp.}} - C_{p,\text{calc.}})$ are large and positive. Up T_s , calculated (with K-N rule) and experimental values are in rather good agreement.

Comparing the $C_p = f(T)$ curves obtained for all these alloys it can be concluded that T_s corresponds to the appearance of an equilibrium in the solid state.

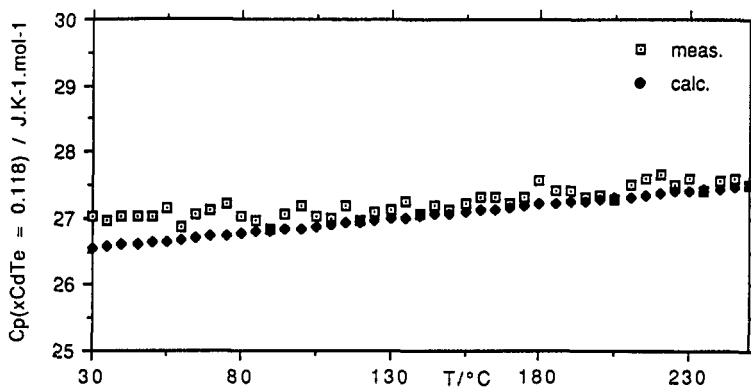


Fig. 1 : Molar heat capacity of CdTe(0.118)HgTe(0.882) solid alloy on the temperature range 30°C-250°C.

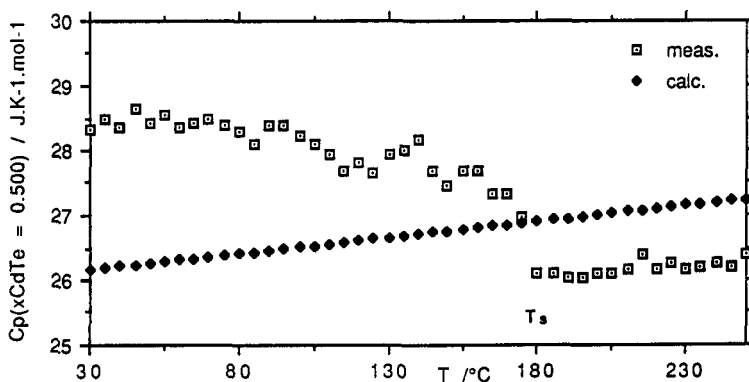


Fig. 2 : Molar heat capacity of CdTe(0.500)HgTe(0.500) solid alloy on the temperature range 30°C-250°C.

The large value of ΔC_p may be due to the presence of two phases and to the thermal effect associated to the change of the equilibrium (solid \leftrightarrow solid (1) + solid (2)) during the heating (from T_1 to T_2) of the C_p measurement. In the one-phase region ($T = 500$ K) between the solidus line and the miscibility gap, ΔC_p is negative corresponding probably to an ordering process on the lattice.

2. Emf measurements

On the whole composition range of the solid $(\text{CdTe})_x(\text{HgTe})_{1-x}$ system, variations of the partial free energy of cadmium telluride were determined by using the potentiometric technique. These measurements were carried out with solid alloys containing an excess of tellurium, by using the electrolytic cells represented diagrammatically thus :

Cell n°1 (-) Cd (sol.) / $\text{CaCl}_2 + \text{CdCl}_2$ in glycerol / $((\text{CdTe})_x(\text{HgTe})_{1-x} + \text{Te})$ (sol.) (+)

Cell n°2 (-) Cd (liq.) / $\text{RbCl} + \text{LiCl} + \text{CdCl}_2$ melt / $((\text{CdTe})_x(\text{HgTe})_{1-x} + \text{Te})$ (sol.) (+)

- The cell n°1 is operable between 307 K and 416 K and the liquid electrolyte is made of pure glycerol saturated in CaCl_2 and containing 0.05 weight per cent of CdCl_2 .

- The cell n°2 is operable between 580 K and 700 K. $\text{RbCl} + \text{LiCl}$ eutectic mixture with 0.05 wt. % of cadmium chloride constitutes the electrolyte.

Reference and working electrodes are made of a tungsten wire attached to a pelletised mixture of metal or alloys in powder form. A special mould allows to press isostatically the powdered compounds intimately mixed around the W wire. This technique insures a good electric contact between electrodes (pellets) and electric leads (W).

All requirements of a potentiometric study concerning the reversibility of the cell and the measurements of emf and temperatures were carefully observed : emf values were carried out at constant temperature with an digital voltmeter with a large input impedance : an emf value was accepted when the difference between two or three measurements was less than ± 0.3 mV. The experimental temperature was obtained at ± 1 K.

Ten alloys ($x_{\text{CdTe}} = 0.06, 0.10, 0.22, 0.30, 0.40, 0.50, 0.60, 0.70, 0.80, 1.0$ with an excess of about 10 at.% Te) were studied by potentiometry.

All the emf experimental values have been reported in Fig. 3 ; solid and dashed lines correspond respectively of measured and interpolated (and extrapolated) values. On the basis of the compiled data (ref.10), the $E = f(T)$ values obtained with the cell n°1 in which liquid cadmium is the reference electrode were recalculated to be referred to the superheated solid cadmium. Emf (experimental and bibliographic (ref. 10)) values of Cd-Te cells are indicated in the upper part of Fig. 3.

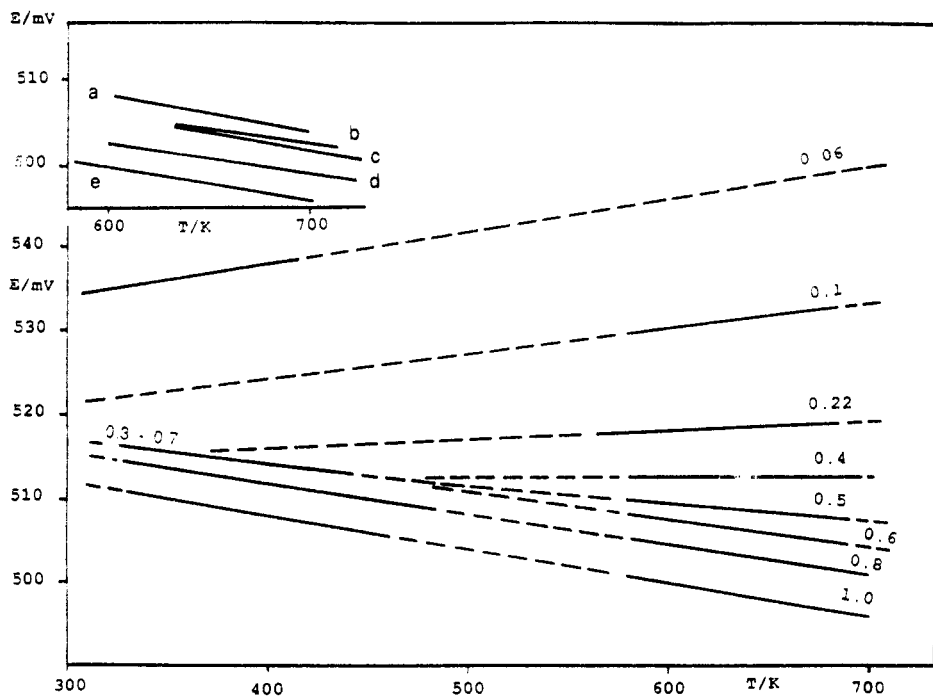


Fig. 3 : $E/\text{mV} = f(T/\text{K})$ for the following molar fractions 0.1, 0.8, 0.6, 0.5, 0.4, 0.22, 0.1, 0.06 . Full and dashed lines correspond to measured and interpolated (or extrapolated) values. In the upper part of the graph, bibliographic (a, b, c, d (ref. 10)) and experimental (e) results are reported.

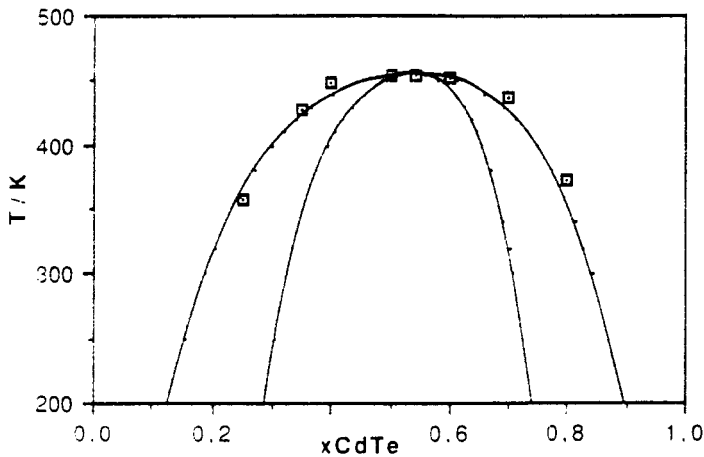


Fig. 4 : Miscibility gap in the solid phase region of the CdTe-HgTe system. Binodal and spinodal curves have been calculated from thermodynamic data (full lines). Experimental points (\square) from calorimetric measurements.

The analysis of the plots $E=f(T)$ leads us to the following remarks :

- over the experimental temperature range, the regular variation of E with the molar fraction x_{CdTe} ascertains the existence of a complete solid solubility at high temperature ;
- the "fan-wise arrangement" of the lines $E = f(T)$ - with a tendency to the union at $T < 500$ K shows :
 - i) the increase of the stability of solid solutions with the temperature,
 - ii) the existence of a heterogeneous region at low temperature (326K - 423 K).

This last remark evidences the independance of emf values with the molar fraction (x_{CdTe}) when $0.3 < x_{CdTe} < 0.7$. In the heterogeneous region of the phase diagram, a least squares fit to the data conformed to a straight line in the temperature range of measurements. These results may be ascribed to the appearance of a miscibility gap ($S_1 + S_2$) at low temperature. Moreover the enthalpies of formation of solid solutions $Cd_xHg_{1-x}Te$ are positive on the molar fraction range $0.1 < x_{CdTe} < 0.9$ (therefore in agreement with the existence of a miscibility gap) and practically equal to zero between 0 - 0.1 and 0.9 - 1.

The positive variation of the free energy of formation with T indicates the decrease of the thermodynamic stability of the solid phase. With all these facts, a solid miscibility gap may be expected in this region under a critical temperature. The limit of this miscibility gap was calculated using the condition that the chemical potentials of components in a two phase region are equal i.e.

$$(\mu_{CdTe})_{(solid\ 1)} = (\mu_{CdTe})_{(solid\ 2)} \quad \text{and} \quad (\mu_{HgTe})_{(solid\ 1)} = (\mu_{HgTe})_{(solid\ 2)}$$

The calculated coordinates of the critical point are: $x_{CdTe} = 0,535$ et $T = 456,5$ K.

Fig.4 illustrates the good agreement between the limit of miscibility gap obtained by potentiometric and calorimetric experiments. The equation of the miscibility gap has been deduced (with $x = x_{CdTe}$) :

$$T/K = - 557.11 + 7216.9 x - 1.960310^4 x^2 + 2.4090 10^4 x^3 - 1.1306 10^4 x^4$$

with $300 < T/K < 455$. The coordinates of the critical point are : $x_{CdTe} = 0.55$ and $T = 455$ K. Taking into account the accuracy of these two series of results, the discrepancy between the coordinates of the critical point are negligible.

In conclusion, determinations of the molar heat capacity ensures the existence of a miscibility gap in the solid CdTe-HgTe system.

Cd-Ga-In-Sn-Zn system

A complete study of the thermodynamic properties of a quinary system (Cd-Ga-In-Sn-Zn) requires theoretically the knowledge of the equivalent properties of 10 binary, 10 ternary and 5 quaternary systems.

The enthalpies of formation of several liquid binary, ternary, quaternary and quinary liquid alloys were determined by direct reaction calorimetry at high temperature on large molar fraction ranges by our group. These measurements were performed using a high temperature Calvet calorimeter or a very high temperature calorimeter equipped with an automatized thermostated charger.

The alloying process was performed in graphite crucible under a high purity argon atmosphere (before entering in the calorimetric cell, the argon flow was passed over titanium sponge heated at 1120 K). The calorimeter was calibrated by Joule effect and by drops of α -alumina (NIST) or by supplementary additions of the first metal.

The uncertainty of the results depends of the number of the components. Errors of the results were estimated of about 1% for a binary system and of about 5 - 7 % for a quinary system.

Previously, the available thermodynamic data of the binary alloys (Cd-Ga, Cd-In, Cd-Sn, Cd-Zn, Ga-In, Ga-Sn, Ga-Zn, In-Sn, In-Zn, Sn-Zn) and of ternary alloys (Ga-In-Sn, Ga-In-Zn, Ga-Sn-Zn, Cd-Ga-In) were analysed. Moreover we have measured the missing enthalpies of formation of Cd(0.33)Ga(0.33)Sn(0.33) and Cd(0.33)Ga(0.33)Zn(0.33) liquid alloys.

The enthalpies of formation of the following liquid quaternary systems have been measured :

Ga-In-Sn-Zn subsystem

- by dropping of Zn metal in the Ga(0.33)In(0.33)Sn(0.33) liquid alloy ;
- by dropping of Sn metal in the Ga(0.33)In(0.33)Zn(0.33) liquid alloy ;
- by dropping of In metal in the Ga(0.33)Sn(0.33)Zn(0.33) liquid alloy .

The experimental values of the integral molar enthalpy of formation at the barycentric point are respectively : 2,269 kJ.mol⁻¹ (addition of Zn), 2,364 kJ.mol⁻¹ (addition of Sn) and 2,343 kJ.mol⁻¹ (addition of In). These results are in a good agreement(mean value $\Delta_{\text{mix}}H_m = 2.325 \text{ kJ.mol}^{-1}$).

Cd-Ga-In-Sn subsystem

- by dropping of Cd metal in the Ga(0.33)In(0.33)Sn(0.33) liquid alloy ;
- by dropping of Sn metal in the Cd(0.33)Ga(0.33)In(0.33) liquid alloy ;
- by dropping of In metal in the Cd(0.33)Ga(0.33)Sn(0.33) liquid alloy.

At the barycentre of this quaternary alloy the experimental values of $\Delta_{\text{mix}}H_m$ are : 1,666 kJ.mol⁻¹ (addition of Cd), 1,965 kJ.mol⁻¹ (addition of Sn) and 2,087 kJ.mol⁻¹ (addition of In). The value obtained by drops of cadmium seems too low. So $\Delta_{\text{mix}}H_m = 2,036 \text{ kJ.mol}^{-1}$ will be used for the calculations.

Cd-Ga-In-Zn subsystem

- by dropping of Cd metal in the Ga(0.33)In(0.33)Zn(0.33) liquid alloy ;
- by dropping of In metal in the Cd(0.33)Ga(0.33)Zn(0.33) liquid alloy ;
- by dropping of Zn metal in the Cd(0.33)Ga(0.33)In(0.33) liquid alloy.

At the barycentric composition, the values of the enthalpy of mixing are : 2,559 kJ.mol⁻¹ (addition of Zn), 2,998 kJ.mol⁻¹ (addition of Cd) and 3,195 kJ.mol⁻¹ (addition of In). $\Delta_{\text{mix}}H_m = 3,096 \text{ kJ.mol}^{-1}$ will be used.

Using the same method, the enthalpies of formation of the Cd-Ga-In-Sn-Zn liquid system have been investigated. Pure cadmium, tin and zinc have been added at 730 K to Ga(0.25)In(0.25)Sn(0.25)Zn(0.25), Cd(0.25)Ga(0.25)In(0.25)Sn(0.25) and Cd(0.25)Ga(0.25)In(0.25)Sn(0.25) quaternary liquid alloys. Experimental results are reported on the graphs (Fig. 5 and 6) and Table 1. At the equimolar composition the enthalpy of formation obtained by dropping cadmium or zinc metal are in good agreement ; the value obtained by adding tin metal is too low.

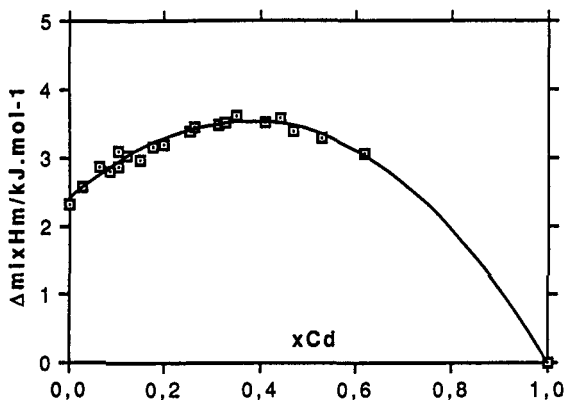


Fig. 5 : Integral molar enthalpy of formation (in kJ.mol⁻¹) of Cd-Ga-In-Sn-Zn system from the Ga(0.25)In(0.25)Sn(0.25)Zn(0.25) liquid alloy at 730 K.

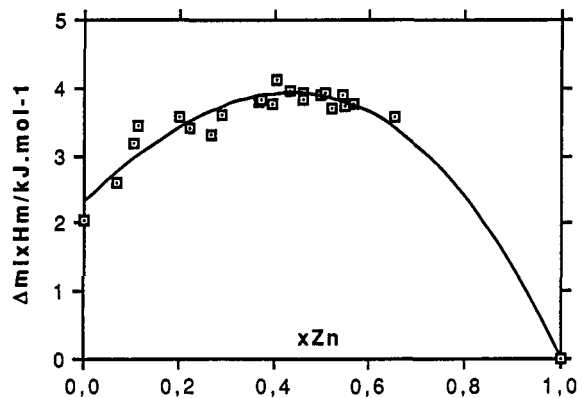


Fig. 6 : Integral molar enthalpy of formation (in kJ.mol⁻¹) of Cd-Ga-In-Sn-Zn system from the Cd(0.25)Ga(0.25)In(0.25)Sn(0.25) liquid alloy at 730 K.

Table 1 : Enthalpies of formation of the liquid Cd-Ga-In-Sn-Zn system at 730 K.

T = 730 K	Ga(0.25)In(0.25)Sn(0.25)Zn(0.25)	T = 730 K	Ga(0.25)In(0.25)Sn(0.25)Zn(0.25)
xCd	$\Delta_{\text{mix}}H_m/\text{J.mol}^{-1}$	xCd	$\Delta_{\text{mix}}H_m/\text{J.mol}^{-1}$
0.000	2325	0.256	3379
0.029	2572	0.265	3443
0.064	2886	0.315	3496
0.088	2812	0.326	3510
0.103	3093	0.352	3628
0.105	2873	0.411	3510
0.123	3042	0.439	3593
0.149	2968	0.469	3383
0.176	3159	0.526	3283
0.200	3188	0.617	3058
T = 730 K	Cd(0.25)Ga(0.25)In(0.25)Zn(0.25)	T = 730 K	Cd(0.25)Ga(0.25)In(0.25)Zn(0.25)
xSn	$\Delta_{\text{mix}}H_m/\text{J.mol}^{-1}$	xSn	$\Delta_{\text{mix}}H_m/\text{J.mol}^{-1}$
0.000	3096	0.320	2524
0.071	3116	0.401	2379
0.136	2691	0.420	2215
0.154	2803	0.461	2271
0.196	2744	0.481	1919
0.253	2563	0.505	2003
0.268	2496	0.510	1761
0.301	2458	0.539	1622
T = 730 K	Cd(0.25)Ga(0.25)In(0.25)Sn(0.25)	T = 730 K	Cd(0.25)Ga(0.25)In(0.25)Sn(0.25)
xZn	$\Delta_{\text{mix}}H_m/\text{J.mol}^{-1}$	xZn	$\Delta_{\text{mix}}H_m/\text{J.mol}^{-1}$
0.000	2036	0.407	4120
0.067	2582	0.435	3971
0.103	3171	0.460	3925
0.115	3452	0.461	3820
0.200	3586	0.498	3904
0.226	3423	0.507	3925
0.269	3301	0.521	3715
0.292	3598	0.543	3887
0.372	3799	0.548	3720
0.375	3841	0.566	3782
0.399	3770	0.654	3568

During the last thirty years, several models (empirical and theoretical models) were proposed to predict the thermodynamic properties of ternary alloys. With empirical models it is well-known that thermodynamic properties of ternary mixtures can be computed from equivalent thermodynamic values of the limiting binary system.

In 1990, for ternary alloys, a general three-factors regular-type model then a general multifactor regular-type model have been presented by Wang and co-workers (ref. 11). These two models reproduced all the well-know special models (Kohler, Bonnier, Muggianu, Toop,...) and several dozen of new ones. In 1991, Silhol (ref.12) in her thesis extended these relations to n-components systems. In the same work, Silhol proposed also an extension of the "surrounded atom model" to quaternary and quinary alloys and compared experimental and predicted results.

If for ternary metallic systems, it is sometimes possible to find a good representation of enthalpy of formation by using an empirical model, for quaternary and quinary systems a satisfactory fit between experimental and predicted values from the same empirical model is very scarce. For instance, experimental results obtained for the quaternary are in good agreement with those obtained by using the Toop model (Fig. 7) ; on the other hand, for the quinary alloy a better representation was obtained with the Hoch model (Fig. 8).

For many other systems (Ag-Au-Si, Ag-Au-Ge, Pd-Ni-Ga, Pd-Ni-In,...), we met the same difficulty to predict the excess function of formation.

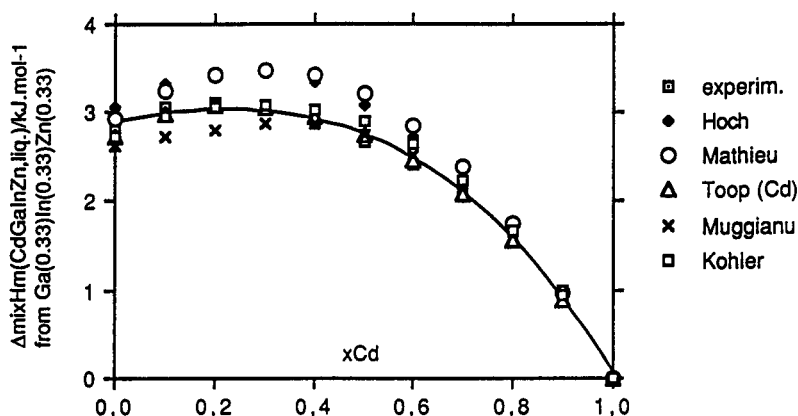


Fig. 7 : Experimental and calculated integral molar enthalpy of formation (in kJ.mol⁻¹) of liquid Cd-Ga-In-Zn system at 730 K.

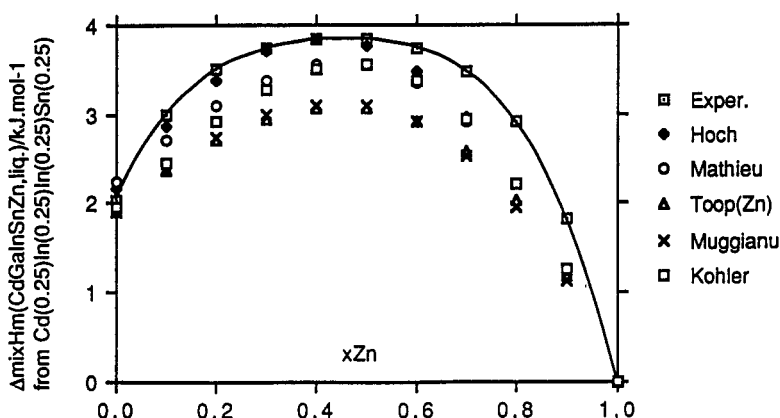


Fig. 8 : Experimental and calculated integral molar enthalpy of formation (in kJ.mol⁻¹) of liquid Cd-Ga-In-Sn-Zn system at 730 K.

CONCLUSION

In materials science, accurate high temperature calorimetric measurements are fairly often unavoidable to understand the behaviour of materials. So it is most important to have the will to develop precise calorimetric techniques to measure enthalpy of formation and molar heat capacity of n-component materials.

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